

Richard Bream,\* David Watkin  
and Andrew Cowley

Chemical Crystallography, Central Chemistry  
Laboratory, University of Oxford, Oxford OX1  
3TA, England

Correspondence e-mail:  
richard.bream@pmb.ox.ac.uk

#### Key indicators

Single-crystal X-ray study  
 $T = 110$  K  
Mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å  
 $R$  factor = 0.070  
 $wR$  factor = 0.097  
Data-to-parameter ratio = 25.6

For details of how these key indicators were  
automatically derived from the article, see  
<http://journals.iucr.org/e>.

## Methylcyclopentane

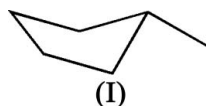
Methylcyclopentane,  $\text{C}_6\text{H}_{12}$ , a liquid at room temperature, was studied as part of a project to develop a computer-controlled low-temperature crystal-growing device. A single crystal was obtained at 115 K. The ring has an envelope conformation, with a pseudo-equatorial methyl substituent on the flap atom.

Received 13 February 2006

Accepted 17 February 2006

### Comment

The melting point of methylcyclopentane is noted by the CRC Handbook of Chemistry and Physics as being  $-142.4^\circ\text{C}$  (130.8 K) (Weast, 1978). A sample solidified spontaneously to a polycrystalline mass on flash-cooling to 115 K, and was then zone refined into a single crystal using tandem computer-controlled heating elements. Data collection was completed at 110 K



The molecule is in the envelope conformation (Fig. 1), with the four atoms C2–C5 almost coplanar (maximum deviation 0.04 Å) and a pseudo-equatorial methyl group attached to the flap atom C1. The crystal structure consists of molecular stacks formed by unit-cell translations along the  $a$  axis (Figs. 2 and 3).

The calculated density is not unlike that of the ordered monoclinic phase of cyclohexane ( $0.996\text{ Mg m}^{-3}$ ), suggesting that a low density may be a feature of small cyclic hydrocarbons (Kahn *et al.*, 1973).

### Experimental

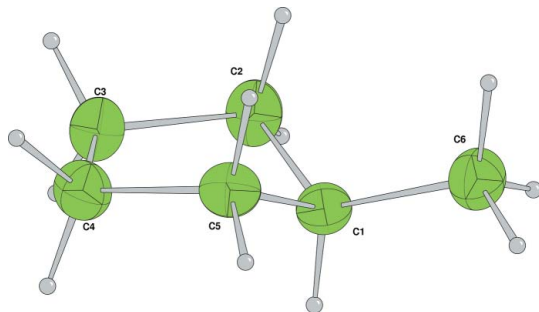
The material was used as supplied by Acros Organics. A 2.0 mm column was flame-sealed in a 0.3 mm diameter Lindemann tube and crystallized as described above.

#### Crystal data

$\text{C}_6\text{H}_{12}$	$D_x = 0.968\text{ Mg m}^{-3}$
$M_r = 84.16$	Mo $K\alpha$ radiation
Monoclinic, $P2_1/n$	Cell parameters from 1447 reflections
$a = 5.3934$ (2) Å	$\theta = 5\text{--}28^\circ$
$b = 11.1439$ (5) Å	$\mu = 0.05\text{ mm}^{-1}$
$c = 9.7047$ (5) Å	$T = 110\text{ K}$
$\beta = 98.0288$ (17)°	Cylinder, colourless
$V = 577.57$ (5) Å <sup>3</sup>	$1.00 \times 0.20$ (radius) mm
$Z = 4$	

#### Data collection

Nonius KappaCCD diffractometer	1412 independent reflections
$\omega$ scans	1408 reflections with $I > 3\sigma(I)$
Absorption correction: multi-scan (DENZO/SCALEPACK; Otwinowski & Minor, 1997)	$R_{\text{int}} = 0.063$
$T_{\text{min}} = 0.69$ , $T_{\text{max}} = 0.98$	$\theta_{\text{max}} = 28.3^\circ$
7788 measured reflections	$h = -7 \rightarrow 7$
	$k = -14 \rightarrow 14$
	$l = -12 \rightarrow 12$



**Figure 1**  
The title compound with displacement ellipsoids drawn at the 50% probability level. H atoms are shown as spheres of arbitrary radius.

**Refinement**

Refinement on  $F^2$   
 $R[F^2 > 2\sigma(F^2)] = 0.070$   
 $wR(F^2) = 0.097$   
 $S = 0.94$   
 1408 reflections  
 55 parameters  
 H-atom parameters constrained

$$w = 1/[\sigma^2(F^2) + (0.03P)^2 + 0.1P]$$

where  $P = [\max(F_o^2, 0) + 2F_c^2]/3$   
 $(\Delta/\sigma)_{\max} < 0.001$   
 $\Delta\rho_{\max} = 0.28 \text{ e } \text{Å}^{-3}$   
 $\Delta\rho_{\min} = -0.22 \text{ e } \text{Å}^{-3}$

**Table 1**  
Selected geometric parameters (Å, °).

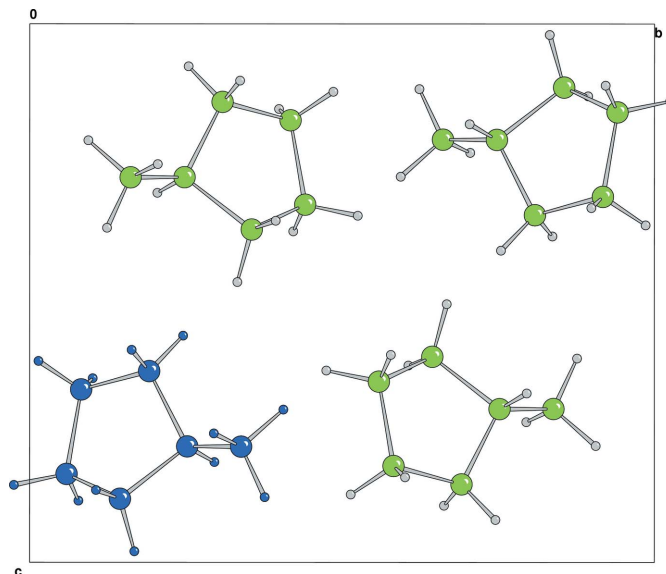
C1—C2	1.5290 (14)	C2—C3	1.5344 (14)
C1—C5	1.5274 (14)	C3—C4	1.5402 (15)
C1—C6	1.5171 (14)	C4—C5	1.5286 (14)
C2—C1—C5	101.82 (8)	C2—C3—C4	105.88 (8)
C2—C1—C6	114.77 (8)	C3—C4—C5	105.39 (8)
C5—C1—C6	115.00 (8)	C4—C5—C1	104.05 (8)
C1—C2—C3	105.28 (8)		

The H atoms were all located in a difference map and then repositioned geometrically. The H atoms were initially refined with soft restraints on the bond lengths and angles to regularize their geometry (C—H in the range 0.93–0.98 Å) and displacement parameters [ $U_{\text{iso}}(\text{H})$  in the range 1.2–1.5 times  $U_{\text{eq}}$  of the parent atom], after which they were refined with riding constraints.

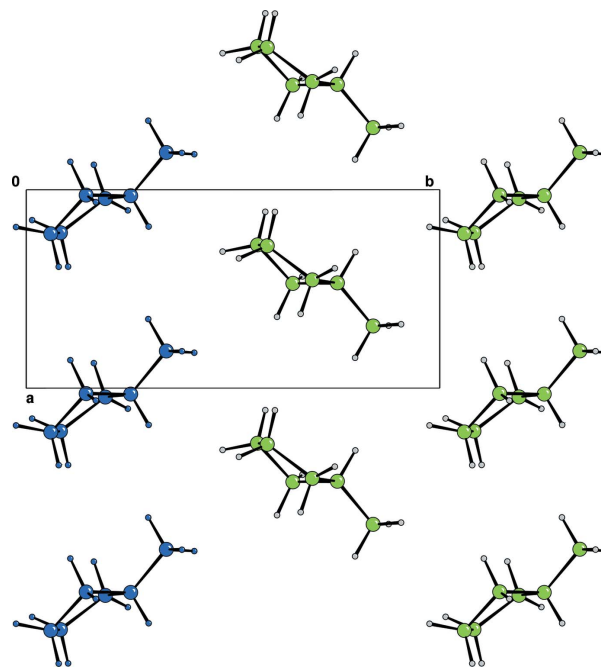
Data collection: *COLLECT* (Nonius, 2001); cell refinement: *DENZO/SCALEPACK*; data reduction: *DENZO/SCALEPACK* (Otwinowski & Minor, 1997); program(s) used to solve structure: *SIR92* (Altomare *et al.*, 1994); program(s) used to refine structure: *CRYSTALS* (Betteridge *et al.*, 2003); molecular graphics: *CAMERON* (Watkin *et al.*, 1996); software used to prepare material for publication: *CRYSTALS*.

**References**

Altomare, A., Cascarano, G., Giacovazzo, G., Guagliardi, A., Burla, M. C., Polidori, G. & Camalli, M. (1994). *J. Appl. Cryst.* **27**, 435.  
 Betteridge, P. W., Carruthers, J. R., Cooper, R. I., Prout, K. & Watkin, D. J. (2003). *J. Appl. Cryst.* **36**, 1487.  
 Kahn, R., Fourme, R., Andre, D. & Renaud, M. (1973). *Acta Cryst.* **B29**, 131–138.  
 Nonius (2001). *COLLECT*. Nonius BV, Delft, The Netherlands.



**Figure 2**  
An *a*-axis projection of the title compound. One column of molecules has been highlighted in blue for comparison with Fig. 3.



**Figure 3**  
A projection along the *c* axis, showing the molecular stacks parallel to the *a* axis.

Otwinowski, Z. & Minor, W. (1997). *Methods in Enzymology*, Vol. 276, *Macromolecular Crystallography*, Part A, edited by C. W. Carter Jr & R. M. Sweet, pp. 307–326. New York: Academic Press.  
 Watkin, D. J., Prout, C. K. & Pearce, L. J. (1996). *CAMERON*. Chemical Crystallography Laboratory, Oxford, England.  
 Weast, R. C. (1978). Editor. *CRC Handbook of Chemistry and Physics*. Cleveland, Ohio: CRC Press.